

of dry solid ammonium bromide. When these reactions were carried out under suitable conditions, the corresponding colorless trisubstituted methane was found to be optically active. In one experiment a specific rotation  $[\alpha]_D^{20}$  in acetone  $-11.7^\circ$  was obtained. Since this methane derivation was obtained in an optically active condition it follows that the colored sodium triarylmethyl must have been active. Hence the three groups and the central carbon atom of the ion could not have been in a plane.

The optical stability of such an ion is being investigated further and a discussion of these results will be published at a later date.

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#### OXIDATION COLORS DERIVED FROM 5,6-DIAMINOURACIL

*Sir:*

Ortho-, meta- and para-diamines differ in their behavior toward oxidizing agents, such as hydrogen peroxide, as follows. With ortho-diamines two molecules react to form a diaminophenazine. With para-diamines two or more molecules react to form a quinone-anil of the type of Bandrowski's base. Meta-diamines, however, do not react. Nevertheless, in the presence of an ortho- or para-diamine (which is capable of forming an ortho- or para-quinoid derivative on oxidation), meta-diamines do react to form deeply colored substances which have found application in fur and hair dyeing.

We find that 5,6-diaminouracil (an ortho-diaminopyrimidine) is also capable of reacting with meta-diamines, aminophenols and diphenols, in the presence of hydrogen peroxide, to form colored substances which dye animal fibers. Thus, for example, the following colors have been obtained from this pyrimidine: with *m*-phenylenediamine, purple; with *m*-aminophenol, old rose; with resorcinol, salmon.

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#### THE CONCENTRATION AT WHICH HEATS OF DILUTION ARE MEASURED IN THE CALORIMETRIC METHOD: A CORRECTION

*Sir:*

In a recent letter [THIS JOURNAL 54, 4114 (1932)] we objected to certain of the claims made by E. Lange and A. L. Robinson based upon the extrapolation of their calorimetric data for intermediate heats of dilution